

Standard: #3-1 & #3-3
Chemistry
Ionic Covalent Mix Nomenclature
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Ionic Compounds
Gives the name not the quantity

Covalent
Always gives the Quantity

	Formula	Type of compound
1. Sulfur dioxide =	<u>SO₂</u>	_____
2. Aluminum chloride =	<u>AlCl₃</u>	_____
3. Copper (II) fluoride =	<u>CuF₂</u>	_____
4. Hydrogen oxide =	<u>H₂O (water)</u>	_____
5. Hydrogen dioxide =	<u>H₂O (water)</u>	_____
6. Pentaphosphorous decoxide =	<u>P₅O₁₀</u>	_____
7. Copper (I) nitride =	<u>Cu₃N</u>	_____
8. Sodium bicarbonate =	<u>NaHCO₃</u>	_____
9. Carbon tetrachloride =	<u>CCl₄</u>	_____
10. Iron (III) hypochlorite =	<u>Fe(ClO)₃</u>	_____
11. Ammonium sulfate =	<u>(NH₄)₂SO₄</u>	_____
12. Gold (IV) oxide =	<u>AuO₂</u>	_____
11. Ammonium Phosphate =	<u>(NH₄)₃PO₄</u>	_____

See
Ion
sheet

	Name	Bond type
1. CuCl =	<u>Copper(I) Chloride</u>	_____
2. CO =	<u>Carbon Monoxide</u>	_____
3. AlCl ₃ =	<u>Aluminum Chloride</u>	_____
4. HClO =	<u>Hydrogen Chlorite</u> / <u>Hydrogen hypochlorite</u>	_____
5. Sn(C ₂ O ₄) ₂ =	<u>Tin(IV) oxalate</u>	_____
6. Zn(OH) ₂ =	<u>Zinc Hydroxide</u>	_____
7. Ca ₃ (PO ₄) ₂ =	<u>Calcium Phosphate</u>	_____
8. PCl ₅ =	<u>Phosphorous pentachloride</u>	_____
9. NH ₄ SCN =	<u>Ammonium thiocyanate</u>	_____
10. P ₂ O ₅ =	<u>diphosphorous pentoxide</u>	_____
11. Al ₂ (SO ₃) ₃ =	<u>Aluminum Sulfite</u>	_____
12. N ₂ O ₅ =	<u>dinitrogen pentoxide</u>	_____
13. (NH ₄) ₃ P =	<u>Ammonium Phosphide</u>	_____